

## Defining Acids + Bases

1) a) both

It can donate  $H^+$  to become  $OH^-$ .  
It can accept  $H^+$  to become  $H_3O^+$ .

b) base

It can accept  $H^+$  to become  $H_2O$ .

c) base

It can accept  $H^+$  to become  $NH_4^+$ .

d) acid

It can donate  $H^+$  to become  $NH_3$ .

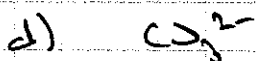
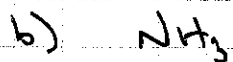
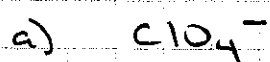
e) base

It can accept  $H^+$  to become  $NH_3$ .

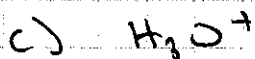
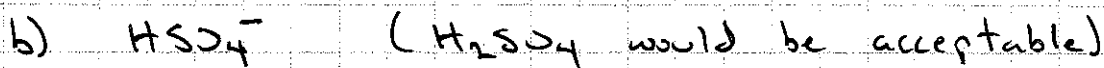
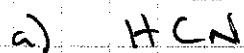
f) base

It can accept  $H^+$  to become  $HCO_3^-$ .

② To find the conjugate base, you simply remove one  $H^+$  from the formula of the acid.



③ To find the conjugate acid, you simply add on  $H^+$  to the formula of the base.



④ To identify each pair, look for a reactant and a product that differ by a single  $H^+$ .

